

ANSWER 1 OF 1 CAPLUS COPYRIGHT 2010 ACS on STN
AN 2000:626261 CAPLUS Full-text
DN 133:302614
ED Entered STN: 10 Sep 2000
TI Electrolysis of brine
CS Akzo Nobel, UK
SO Research Disclosure (2000), 436(Aug.), P1408 (Number 436093)
CODEN: RSDSBB; ISSN: 0374-4353
PB Kenneth Mason Publications Ltd.
DT Journal; Patent
LA English
CC 72-9 (Electrochemistry)
Section cross-reference(s): 48
FAN.CNT 1

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
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PI RD 436093 20000810 <--		20000810	RD 2000-436093	
PRAI RD 2000-436093		20000810		

AB It is fount that formic acid, oxalic acid and tartaric acid can be used as alternative to hydrogen peroxide and sodium bisulfate for removal of active chlorine from brine in process of its electrolysis.

ANSWER 1 OF 1 WPIX COPYRIGHT 2010 THOMSON REUTERS on STN
AN 2001-078626 [200109] WPIX Full-text
DNC C2001-022406 [200109]
DNN N2001-059813 [200109]
TI Electrolysis of brine involves adding formic acid, oxalic acid and tartaric acid to remove active chlorine from brine
DC E34; E36; J03; X25
PA (ALKU-C) AKZO NOBEL
CYC 1
PI RD 436093 A 20000810 (200109)* EN 0[0]
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ADT RD 2000-436093 20000720
PRAI RD 2000-436093 20000720
IC ICM C25B0000-00
AB RD 436093 A UPAB: 20050524
NOVELTY - Formic acid, oxalic acid and tartaric acid are used as alternatives to, e.g., hydrogen peroxide and sodium bisulfite for the removal of active chlorine from brine.
USE - The method is used for membrane chloro-alkali plants utilizing a re-circulating brine system.
ADVANTAGE - By using formic acid, oxalic acid and tartaric acid, no undesired by-products are formed that can decompose free chlorine. The only reaction products are chloride, carbon dioxide and water.
FS CPI; EPI
MC CPI: E10-C02D1; E10-C02E; E10-C04J1; E31-B01; E33-A01; J03-B04
EPI: X25-R01